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Information . Molar gas volume = 24.0 dm³ mol⁻¹ at room temperature and pressure, RTP Avogadro constant, $N_A = 6.02 \times 10^{23} \text{ mol}^{-1}$ Specific heat capacity of water, $c = 4.18 \text{ J g}^{-1} \text{ K}^{-1}$ Ionic product of water, $K_w = 1.00 \times 10^{-14} \text{ mol}^2 \text{ dm}^{-6}$ at 298 K 1 tonne = 10⁶ g Arrhenius equation: $k = Ae^{-E_a/RT}$ or $\ln k = -E$

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